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The Crystal Structure of Protopine, C₂₀H₁₉O₅N

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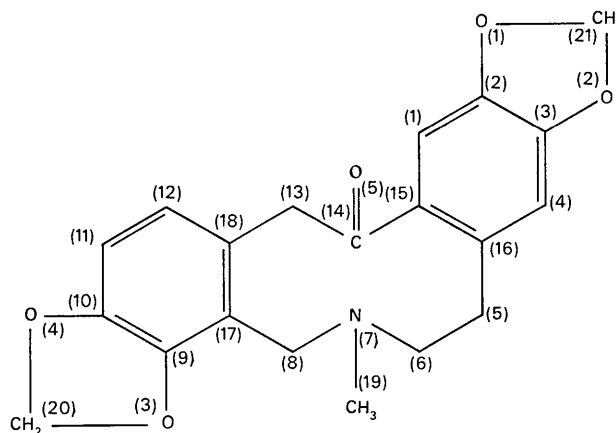
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The crystal structure of protopine, C₂₀H₁₉O₅N, has been determined by the symbolic addition procedure, and has been refined by least squares to an *R* index of 0.043, with use of counter data for 2904 observed reflexions. The space group is *P2₁/n* and the cell constants are *a* = 13.610, *b* = 17.445, *c* = 7.102 Å, β = 96.75°. The ten-membered ring in the molecule is severely buckled and contains a short distance across the ring of only 2.555 Å between the C atom of the carbonyl group and the tertiary nitrogen atom. Each of the two benzene rings has a very slight boat form, and the normals to their mean planes are at an angle of 27.8° from each other. The carbonyl group is about 37.4° out of the mean plane of its nearest benzene ring. The carbonyl oxygen atom and the methyl group are on the same side of the ten-membered ring.

Introduction

A natural group of alkaloids in opium is presented by the two members: protopine, C₂₀H₁₉O₅N, and cryptopine, C₂₁H₂₃O₅N. Protopine is almost the typical alkaloid of the poppy order (Papaveraceae) since it has been found in most species of this order. Molecules in this group are characterized by the presence of a ten-membered *N*-hetero-ring containing one carbonyl group (Manske & Holmes, 1954). The carbonyl group and the tertiary amino function are held very closely together across this medium-sized ring, and it is of interest to define the degree of interaction between them. The two functional groups have been found to interact when a proton is added to the carbonyl oxygen, as for example in protopine hydriodide, with the net result of transforming the ten-membered ring into two six-membered rings. This interaction was studied by Mottus, Schwarz & Marion (1953), and by Anet & Marion (1954) from the ultraviolet and infrared spectra.

The molecular structure for the protopine base is shown in (I); the numbering of the atoms is based on that given by Henry (1949), p. 300.



(I) Protopine

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X-ray analysis of protopine and cryptopine has been undertaken in order to determine their structures and

in particular the shape and dimensions of the central ten-membered ring, and the orientation of the carbonyl group relative to the rings. In each case, the analysis has been carried out on the base, since the heavy atom derivatives have been expected to involve steric change or complete closure of the ten-membered ring. In addition to the chemical interest in the two structures, they have been considered by the authors to be very suitable for trying out a self-consistent set of crystallographic programs developed by Ahmed, Hall, Pippy & Huber (1966) in FORTRAN IV for use on the IBM 360 system, and in particular a program for the direct method of phase determination by the symbolic addition procedure for centrosymmetric crystals. The structure of protopine is reported in this paper and that of cryptopine will be described in a separate communication.

Crystal data

Space group and unit-cell data for protopine, based on precession photographs, have been reported by Barnes & Forsyth (1954) as follows: space group $P2_1/c$; $a=7.109$, $b=17.44$, $c=16.08$ Å, $\beta=122^\circ 45'$, $Z=4$; $D_c=1.399$ g.cm⁻³, $D_o=1.393$ g.cm⁻³.

For the present X-ray analysis the space group has been changed to $P2_1/n$, and the cell constants as re-measured on the General Electric XRD 5 diffractometer using a 1° take-off angle, 0.02° and 0.05° slits and Cu radiation (λ , $K\alpha_1=1.54050$ Å, $K\alpha_2=1.54434$ Å) are $a=13.610$ ($\sigma=0.003$), $b=17.445$ ($\sigma=0.004$), $c=7.102$ Å ($\sigma=0.003$), $\beta=96^\circ 45'$ ($\sigma=3'$); $U=1674.5$ Å³, mol. wt.=353.36, $D_c=1.401$ g.cm⁻³; $\mu(\text{Cu})=8.48$ cm⁻¹, $F(000)=744$. For this space group, the systematic absences are the $h0l$ with $h+l=2n+1$, and the $0k0$ when k is odd.

Intensity data

Cu radiation, Ni filters, and a prismatic crystal of dimensions between 0.2 and 0.3 mm were employed for the intensity data collection by the moving-crystal moving-counter method (Furnas, 1957). The crystal was mounted with the b axis along the direction of the fibre, the take-off angle was set at 2° , and the integrated intensity was measured at each lattice site within the range of the instrument. The 2θ scan was varied between 2° for the low-angle reflexions and 4° for the high-angle reflexions, and the background was measured at the beginning and at the end of each scan on the Laue streak through the reflexion. The mA setting was reduced for the strong reflexions instead of using attenuators, and when the peak height exceeded 20000 counts per second the intensity was corrected for non-linearity of the counter by an empirically evaluated function. The intensity of the 400 reflexion was measured at short intervals of time during the data collection and employed for scaling of the intensities. Many of the weak reflexions were remeasured prior to the final stages of refinement in order to improve their

counting statistics. The $1/L_p$ corrections were applied to the net counts (total - background), and the absorption corrections were ignored since $(\mu t)_{av}=0.21$ (Jeffrey & Rose, 1964).

In the reciprocal space corresponding to $2\theta_{max}$ of about 161° ($\sin \theta/\lambda=0.64$), there are 3705 non-equivalent lattice sites excluding those prohibited by the space group, and of these, 2904 reflexions (78.4%) were observed above threshold.

Structure determination by symbolic addition

The structure was determined by the symbolic addition procedure for centrosymmetric crystals (Hauptman & Karle, 1953), using a program written by S.R.H. for the IBM 360 system in FORTRAN IV, which performs all the necessary operations automatically. The program starts with the $|F_o|$ data on some arbitrary scale and the $f(hkl)$ values at zero temperature, as output from the data reduction program, and concludes by supplying the signed $E(hkl)$ data in the appropriate order on magnetic tape which can be used directly as input to the Fourier program for calculation of the E map. However, human intervention is made possible at various stages in the program if need be.

The scale of the $|F_o|$ data, K' , and the overall temperature factor, B , were estimated from the Wilson plot. In this calculation all the non-equivalent observed and accidentally unobserved reflexions were included, and appropriate allowances were made for the equivalent reflexions and for the systematic absences. The scattering factor curves employed in this and all other calculations were those derived by Hanson, Herman, Lea & Skillman (1964), except that the hydrogen curve was modified by a temperature factor of -2.5 Å² as suggested by Jensen (1965). The normalized structure factor amplitudes $|E_h|$ were calculated from the expression

$$|E_h|^2 = K(s)|F_h|^2/\varepsilon \sum_j f_j^2(s)$$

where $|F_h|$ are on the assumed original scale, $s=\sin \theta/\lambda$, $K(s)=K'^2 \cdot \exp(2Bs^2)$, and ε for space group $P2_1/n$ was taken as 2 for the $h0l$ and $0k0$ reflexions and as 1 for all the others. The computed values for the distribution and statistics of $|E|$ for this structure, and the corresponding theoretical values which were calculated by Karle, Hauptman, Karle & Wing (1958), given in parentheses, were as follows: $\langle |E| \rangle = 0.803$ (0.798), $\langle |E|^2 \rangle = 0.989$ (1.000), $\langle |E^2 - 1| \rangle = 0.954$ (0.968); reflexions with $|E| > 3$ were 0.30% (0.3%), reflexions with $|E| > 2$ were 4.67% (5.0%), and reflexions with $|E| > 1$ were 30.4% (32.0%).

The 478 reflexions with $|E| > 1.50$ were then separated from the rest of the data, sorted in descending order of $|E|$, and each of them was assigned one of the eight parities in order to simplify the search operation. The search for sets of reflexions which satisfy the Σ_2 relationship

$$sE_h \sim s \sum_h E_{h'} \cdot E_{h-h'}$$

was performed for all values of $|E_h| > 1.50$ but limiting $|E_{h'}|$ and $|E_{h-h'}|$ to values more than 2.0 each. The three origin-defining reflexions 874, 735, and 12,6,1 were selected by the program from the first ten reflexions with the largest $|E|$'s, and it was found necessary to select the 1,14,4 reflexion to define a symbol A . The first stage of sign determination was limited to the 178 reflexions with $|E| > 2.0$, but it was possible to fix the signs of only 164 of these reflexions, and it was found later that 14 of them were wrongly signed. The sign determination was then extended to include the remainder of the 478 reflexions with $|E| > 1.50$. Altogether, the signs were deduced for a total of 453 reflexions, of which 47 signs were later found to be wrong. However, the E map which was evaluated with these 453 reflexions gave the structure (excluding the hydrogen atoms) without any ambiguity as shown in Fig. 1, and there were no spurious peaks in the map as high as those selected for the molecule. The relatively high percentage of incorrect signs can be attributed to the unsuitability of the scheme adopted in the program for acceptance or rejection of the signs, and steps are now being taken to modify this part of the program.

The structure factors calculated for the trial model, assuming isotropic temperature factors, gave an R index of 0.27.

Refinement

The structure was refined by successive cycles of block-diagonal least squares, employing only the observed

reflexions and evaluating matrices of sizes 4×4 , 9×9 , or $(3 \times 3$ and $6 \times 6)$ per atom. After two cycles of isotropic refinement, the R index was reduced to 0.14, and after three more cycles of anisotropic refinement, it was reduced further to 0.09. A difference map was computed and positions of all the hydrogen atoms were derived from it. In cycles six to nine, the C, N, and O atoms were refined anisotropically and the H atoms isotropically, and the R index was brought to 0.05. It was obvious at this stage that the structure amplitudes for some of the weak reflexions were not as accurate as the rest of the data, and the very strong reflexions were affected by extinction. The intensities of most of the weak reflexions were remeasured and their structure amplitudes were modified accordingly. The structure factors were recalculated without the contributions of the H atoms giving an R index of 0.074, a second difference map was evaluated omitting the 101, 221, 310, 320, and 321 reflexions which had appreciable extinction errors, and new positions for the hydrogen atoms were derived from it. A composite drawing of this difference map is presented in Fig. 2. The peak heights for the hydrogen atoms were between 0.4 and 0.7 $e \cdot \text{\AA}^{-3}$, the residual electron densities on some of the bonds were as high as 0.2 or 0.25 $e \cdot \text{\AA}^{-3}$, and the background elsewhere was quite flat and within $\pm 0.15 e \cdot \text{\AA}^{-3}$.

The final three cycles of refinement were carried out on all atoms, with the revised data but excluding the five strong reflexions with appreciable extinction errors. The final shifts in the parameters were much smaller than their e.s.d.'s, the final R index for all the observed

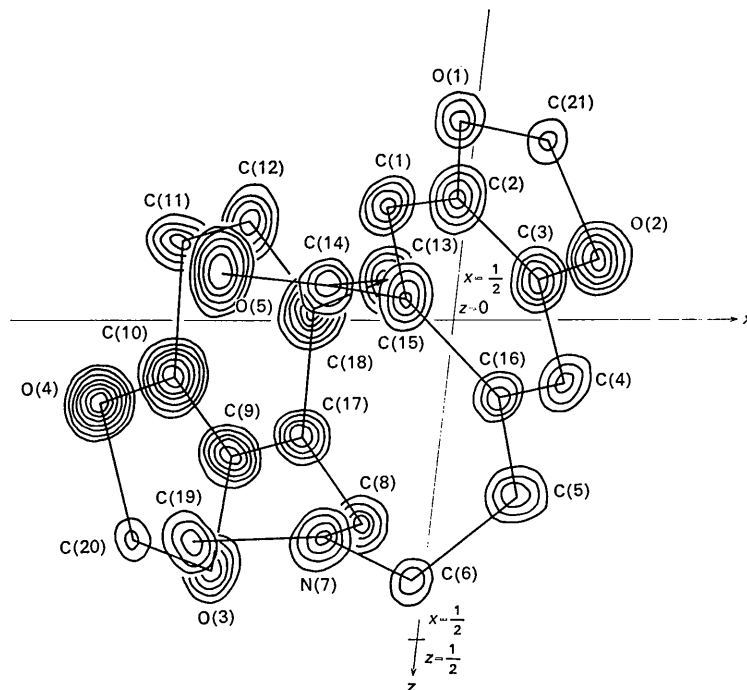


Fig. 1. Composite drawing of the E map as evaluated with 453 reflexions, of which 47 had wrong signs.

data was 0.043, and it was only 0.041 without the five aforementioned reflexions. The $|F_o|$ values for the large majority of the unobserved reflexions were below the threshold amplitudes, and the rest were only very slightly above them. The weights employed in the refinement were calculated by the expression

$$w = 1 / \{1 + [(|F_o| - 60) / 40]^2\}.$$

Results

The refined parameters of the C, N, and O atoms are listed in Table 1, and those of the H atoms are given in Table 2. The e.s.d.'s in Tables 1 and 2 were calculated with expression 6.4.3(6) of *International Tables*

for X-ray Crystallography (1959, p. 330), and increased by 30% in order to allow for omission of the unobserved reflexions and of the interaction between the atoms. The molecule in perspective as viewed from two directions at right angles to each other is shown in Fig. 3.

The observed structure amplitudes, and the calculated structure factors based on the parameters in Tables 1 and 2, are listed in Table 3. A summary of the agreement between $|F_o|$ and $|F_c|$ (Ahmed & Barnes, 1963) is presented in Table 4. Of the 3705 reflexions, only two observed reflexions show high discrepancy; both of these are very strong and are obviously affected by extinction.

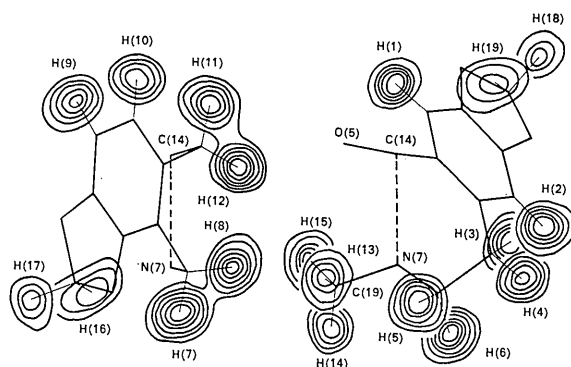


Fig. 2. The second difference map, showing the hydrogen atoms, with the contour lines at intervals of $0.1 \text{ e.}\text{\AA}^{-3}$, starting at $0.2 \text{ e.}\text{\AA}^{-3}$. For clarity, the molecule has been divided at the dotted line.

Table 1. Fractional coordinates, vibration tensor components (\AA^2) for the expression $T = \exp[-2\pi^2(U_{11}a^*2h^2 + \dots + 2U_{23}b^*c^*kl + \dots)]$, and their e.s.d.'s (all quantities $\times 10^4$)

	x	y	z	U_{11}	U_{22}	U_{33}	$2U_{23}$	$2U_{13}$	$2U_{12}$
C(1)	4352 (2)	3173 (1)	-1807 (3)	335 (12)	361 (14)	357 (13)	87 (20)	67 (20)	88 (20)
C(2)	4930 (2)	3808 (1)	-1872 (4)	402 (14)	412 (14)	425 (14)	283 (23)	163 (23)	63 (23)
C(3)	5712 (2)	3951 (1)	-508 (4)	367 (13)	398 (14)	546 (16)	241 (24)	150 (24)	-152 (23)
C(4)	5964 (2)	3462 (1)	974 (4)	339 (13)	427 (14)	443 (14)	166 (23)	-7 (22)	-152 (22)
C(5)	5709 (2)	2263 (1)	2749 (3)	334 (12)	399 (14)	384 (13)	189 (22)	-114 (21)	-118 (22)
C(6)	4882 (2)	2108 (1)	3971 (3)	483 (14)	330 (12)	298 (12)	-4 (20)	3 (21)	-208 (23)
N(7)	4137 (1)	1587 (1)	3037 (2)	316 (10)	259 (10)	303 (9)	-3 (15)	68 (15)	-6 (16)
C(8)	4441 (2)	781 (1)	3222 (3)	357 (12)	276 (12)	294 (11)	34 (19)	-6 (19)	6 (19)
C(9)	3313 (2)	-326 (1)	2099 (3)	338 (12)	274 (12)	376 (13)	42 (19)	88 (20)	64 (19)
C(10)	2738 (2)	-731 (1)	699 (4)	341 (12)	267 (12)	490 (14)	-11 (21)	117 (21)	-51 (20)
C(11)	2682 (2)	-535 (1)	-1168 (4)	432 (14)	373 (14)	429 (14)	-186 (23)	20 (23)	-122 (23)
C(12)	3248 (2)	94 (1)	-1611 (3)	423 (13)	381 (14)	322 (12)	-99 (21)	75 (21)	-38 (22)
C(13)	4433 (2)	1178 (1)	-738 (3)	341 (12)	313 (12)	295 (11)	-6 (19)	137 (19)	6 (20)
C(14)	3945 (2)	1948 (1)	-473 (3)	304 (11)	333 (12)	275 (11)	44 (19)	7 (18)	7 (19)
C(15)	4603 (2)	2644 (1)	-313 (3)	297 (11)	278 (12)	329 (12)	39 (19)	90 (19)	51 (19)
C(16)	5406 (2)	2782 (1)	1062 (3)	280 (11)	331 (12)	354 (12)	85 (19)	69 (19)	-21 (19)
C(17)	3858 (2)	303 (1)	1711 (3)	308 (11)	234 (10)	317 (11)	-39 (18)	59 (18)	53 (19)
C(18)	3830 (2)	505 (1)	-216 (3)	317 (12)	267 (10)	340 (12)	-52 (19)	142 (19)	54 (19)
C(19)	3158 (2)	1697 (1)	3636 (4)	413 (14)	365 (14)	474 (15)	-6 (23)	259 (23)	75 (23)
C(20)	2574 (2)	-1283 (2)	3487 (4)	754 (22)	566 (18)	642 (20)	498 (32)	-336 (34)	-621 (34)
C(21)	5682 (3)	4875 (2)	-2694 (5)	692 (22)	646 (22)	840 (24)	693 (38)	15 (37)	-364 (36)
O(1)	4843 (1)	4391 (1)	-3201 (3)	584 (12)	593 (12)	659 (13)	707 (23)	-45 (21)	-220 (20)
O(2)	6151 (2)	4635 (1)	-886 (3)	614 (13)	555 (12)	802 (15)	641 (23)	-119 (23)	-460 (22)
O(3)	3227 (1)	-651 (1)	3848 (2)	587 (12)	387 (10)	421 (10)	173 (16)	57 (17)	-284 (19)
O(4)	2246 (1)	-1316 (1)	1517 (3)	609 (12)	409 (10)	566 (12)	88 (18)	27 (19)	-418 (19)
O(5)	3052 (1)	2023 (1)	-805 (2)	296 (8)	413 (10)	535 (10)	80 (17)	-121 (15)	25 (16)

The bond lengths and their e.s.d.'s as calculated by the expression of Ahmed & Cruickshank (1953) are shown in Fig. 4(a). The bond angles and their e.s.d.'s (*International Tables for X-ray Crystallography*, 1959, p. 331) are given in Fig. 4(b). These results are based on the refined coordinates and do not include the corrections for thermal vibration. The angles which have been omitted from Fig. 4(b) for simplicity, and their e.s.d.'s, are: C(16)–C(5)–H(4), 106.9° (1.4); C(6)–C(5)–H(3), 110.5° (1.4); N(7)–C(6)–H(5), 105.9° (1.4); C(5)–C(6)–H(6), 108.6° (1.3); N(7)–C(8)–H(8), 110.0° (1.3); C(17)–C(8)–H(7), 109.6° (1.3); C(18)–C(13)–H(12), 111.5° (1.3); C(14)–C(13)–H(11), 102.1° (1.5); N(7)–C(19)–H(14), 111.8° (1.6); H(13)–C(19)–H(15), 105.3° (2.1); O(3)–C(20)–H(17), 115.5° (2.0); O(4)–C(20)–H(16), 106.7° (1.8); O(1)–C(21)–H(19), 112.4° (2.1); O(2)–C(21)–H(18), 103.9° (1.8).

Table 2. Fractional coordinates ($\times 10^3$) and isotropic temperature factors (\AA^2) of the hydrogen atoms, and their e.s.d.'s

	x	y	z	B
H(1)	375 (2)	304 (1)	-276 (3)	3.4 (0.5)
H(2)	655 (2)	359 (1)	194 (3)	3.1 (0.5)
H(3)	600 (2)	175 (1)	231 (3)	3.6 (0.6)
H(4)	629 (2)	252 (1)	353 (3)	3.6 (0.5)
H(5)	451 (2)	260 (1)	425 (3)	3.6 (0.6)
H(6)	519 (2)	187 (1)	527 (3)	3.4 (0.5)
H(7)	436 (2)	56 (1)	450 (3)	3.1 (0.5)
H(8)	520 (2)	73 (1)	308 (3)	3.0 (0.5)
H(9)	224 (2)	-82 (1)	-218 (4)	4.4 (0.6)
H(10)	320 (2)	26 (1)	-297 (4)	3.9 (0.6)
H(11)	448 (2)	117 (1)	-213 (4)	3.9 (0.6)
H(12)	511 (2)	117 (1)	-4 (3)	2.6 (0.5)
H(13)	293 (2)	227 (2)	343 (4)	4.9 (0.7)
H(14)	314 (2)	155 (2)	506 (4)	6.0 (0.7)
H(15)	264 (2)	138 (2)	285 (4)	4.8 (0.6)
H(16)	292 (2)	-180 (2)	384 (4)	7.0 (0.9)
H(17)	196 (3)	-127 (2)	417 (5)	8.6 (1.0)
H(18)	619 (2)	471 (2)	-365 (5)	7.0 (0.9)
H(19)	549 (3)	543 (2)	-259 (5)	9.3 (1.0)

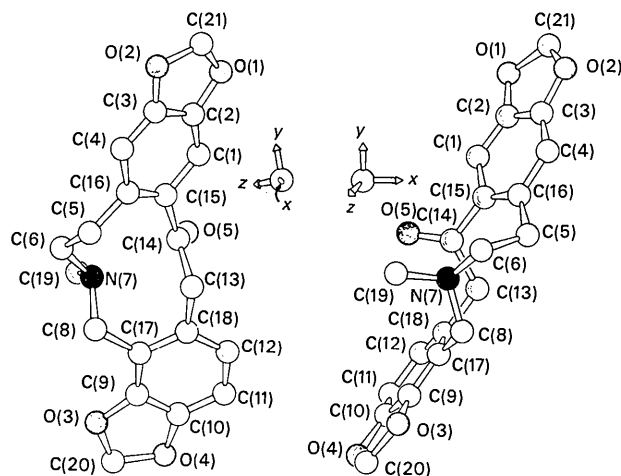


Fig. 3. Two perspective views of the molecule at 90° from each other.

The distance N(7)···C(14) across the ten-membered ring is only 2.555 Å, and it makes angles of 106.0°, 108.6°, and 105.1° with the three bonds N(7)–C(6), N(7)–C(8), and N(7)–C(19), respectively. It also makes angles of 84.7°, 97.7°, and 101.6° with the three bonds C(14)–C(13), C(14)–C(15), and C(14)–O(5), respectively.

Discussion

Bond lengths and angles

The various types of bond in this structure have mean values which are about the same as the corresponding average values given by Sutton (1965), which are quoted here in parentheses. The aromatic C–C bonds in the two benzene rings vary in length from 1.363 to 1.418 Å; the mean is 1.388 Å (1.394 ± 0.005 Å). The C–C single bonds in the ten-membered ring are in the range 1.504 to 1.525 Å, mean is 1.512 Å (1.506 ± 0.005 Å for a C–C single bond shortened in the presence of a carbon–oxygen double bond, and 1.53 ± 0.01 Å in the presence of an aromatic bond). The three C–N bonds are between 1.458 and 1.466 Å; the mean is 1.462 Å (1.472 ± 0.005 Å for a trivalent nitrogen atom). The carbon–oxygen double bond is 1.218 Å, which indicates clearly that it has ketonic character for which the average distance is 1.215 ± 0.005 Å, as compared with the average distance of 1.145 ± 0.01 Å for a carbonyl group (Sutton, 1965). The C–O single bonds are of two types; the first type involves the C-atom of an aromatic ring and varies in length between 1.375 and 1.385 Å, mean is 1.381 Å (1.371 ± 0.016 Å in conjugated heterocyclic compounds), and the second type involves the C-atom of a CH₂ group and ranges from 1.419 to 1.431 Å, mean is 1.425 Å (1.426 ± 0.005 Å in saturated heterocyclic compounds). Finally, the C–H bonds are between 1.00 and 1.06 Å, mean is 1.02 Å (1.06 to 1.09 Å).

The mean angle at each of the twelve carbon atoms in the benzene rings is between 119.98 and 120.00°, thus showing that the three bonds at each of these atoms are nearly planar. At C(14), the mean angle is only 119.3°, and that atom is about 0.11 Å away from the plane through atoms C(13), C(15), and O(5). The four bonds at each of the other carbon atoms are very nearly tetrahedrally arranged; mean angle at each atom is between 109.3 and 109.4°. The three bonds at N(7) and the short distance N(7)···C(14) across the ten-membered ring are arranged nearly tetrahedrally around the nitrogen atom; these angles range from 105.1 to 113.2° and their mean is 109.4°. Therefore, it seems that in derivatives of protopine in which the nitrogen atom is tetravalent, the direction of the fourth bond for least steric change would be across the ring in the direction of N(7)···C(14), thus transforming it into two six-membered rings. The other alternative, in which a fourth atom other than C(14) is attached to N(7), would certainly require a conformational change in the ten-membered ring and a considerable increase in the N(7)···C(14) distance. It is of interest to men-

Table 3. Observed structure amplitudes and calculated structure factors (x 10)

(* indicates unobserved reflexion and |F_{th}| in place of |F_o|).

Table with multiple columns for h, k, l, F_o, F_c, and various reflections. The table lists observed structure amplitudes (F_o) and calculated structure factors (F_c) for various reflections (hkl). The data is organized into several columns, each representing a different set of reflections. The values are given in units of 10. Some reflections are marked with an asterisk (*), indicating they are unobserved. The table includes a wide range of reflections, from low-order to high-order, covering various h, k, and l values.

Table 3 (cont.)

A large grid of numerical data with multiple columns and rows, organized into sections by headers like 'K FD FC' and 'L'. The data consists of integers and some floating-point values, arranged in a structured, repeating pattern across the page.

Table 3 (cont.)

K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC	K	F ₀	FC																																																																			
H=	12	L=	2	4	130	142	1	43	44	5	62	-86	11	36	-31	7	49	36	1	35	34	10	49	-49	1	196	-199	10	200	-11	4	230	-1	0	88	-88	0	100	-100	11	44	-31	5	45	-21	1	120	125	2	39	43	2	30	44	4	31	-74	6	53	-49	7	199	11	8	100	-16	1	25	24	2	32	-38	4	61	57	5	48	-39	6	43	49	7	43	-43	0	90	-82	1	75	74	2	26	24	3	170	3	4	21	14	5	42	40	4	21	14	5	42	40

Table 4. Agreement summary

2904 observed reflexions ($1.4 \leq |F_0| \leq 143.8$)

$$R = 0.043$$

Category	Limits	Number
1	$ \Delta F \leq 1.0 F_{th} $, or $ \Delta F / F_0 \leq 0.10$	2898
2	$1.0 F_{th} < \Delta F \leq 2.0 F_{th} $, or $0.10 < \Delta F / F_0 \leq 0.15$	4
3	$2.0 F_{th} < \Delta F \leq 3.0 F_{th} $, or $0.15 < \Delta F / F_0 \leq 0.20$	1
4	$3.0 F_{th} < \Delta F $, or $0.20 < \Delta F / F_0 \leq 0.30$	1

801 unobserved reflexions ($|F_{cmax}| = 4:1$)

1	$ F_c \leq 1.0 F_{th} $	767
2	$1.0 F_{th} < F_c \leq 1.5 F_{th} $	34

 $|F_{th}|$ = threshold amplitude = 1.3 to 3.4

tion that only a small residual electron density (~ 0.14 e.Å⁻³) was observed on N(7)···C(14) at about 0.75 Å from N(7).

Intermolecular distances

The shortest intermolecular contacts for the different types of atom as found in this structure, and the corresponding distances as calculated from the van der Waals radii (Pauling, 1960, p.260), given in parentheses, are as follows: H···H, 2.40 (2.4); H···O, 2.49 (2.6); C···O, 3.19 (3.1); and C···C, 3.47 (3.4) Å. The nitrogen atom is involved in only two intermolecular distances below 4.0 Å, and both are to H-atoms which are at a distance of 3.48 and 3.99 Å from it. This confirms the proposal by Mottus, Schwarz & Marion (1953), that the nitrogen atom is well buried in the molecule.

Planarity of the rings

The mean plane of the benzene ring containing atom C(1), and of that containing atom C(11), referred to

the orthogonal set of axes $X' = ax + cz \cdot \cos \beta$, $Y' = by$, $Z' = cz \cdot \sin \beta$, are

$$0.6772X' - 0.4862Y' - 0.5523Z' - 2.1388 = 0, \quad (1)$$

and

$$0.7924X' - 0.6018Y' - 0.0999Z' - 3.6218 = 0, \quad (2)$$

respectively. Each of these two rings is non-planar since $\chi^2 = 97$ and 72, respectively, and for $\nu = 3$, $P \ll 0.1\%$. Each ring has a very slight boat form with maximum deviation from the mean plane of only ± 0.013 Å. The angle between the normals to these two planes is 27.8° .

The carbonyl group has atom C(14) about 0.06 Å above plane (1) and atom O(5) about 0.68 Å below it, and the C(14)–O(5) bond at an angle of 37.4° from that plane. This finding is different from the prediction by Mottus, Schwarz & Marion (1953), that the carbonyl is very nearly in plane (1). For the sake of completeness, it should also be mentioned that the carbonyl is at an angle of 35.0° from plane (2).

The ten-membered ring can be described as having the four atoms C(14), C(15), C(16), C(5) approximately

in plane (1), the four atoms C(8), C(17), C(18), C(13) approximately in plane (2), and the two parts are linked together by the bond C(15)–C(16) on one side, and by the three bonds C(5)–C(6), C(6)–N(7), and N(7)–C(8) on the other side. The ring is severely buckled as shown in the two views of the molecule in Fig. 3 and contains a short distance across the ring of only 2.555 Å between N(7) and C(14). Both the N(7)–C(19) and C(14)–O(5) bonds are on the same side of the ring,

the angle between them is 27.9°, and the distance C(19)···O(5) is 3.19 Å. These two bonds are slightly skew, but their mean plane makes exactly equal angles of 60.7° with planes (1) and (2).

Finally, as is to be expected, each of the two five-membered rings is non-planar.

The authors are indebted to Dr W.H. Barnes for suggesting the problem, and for making his original

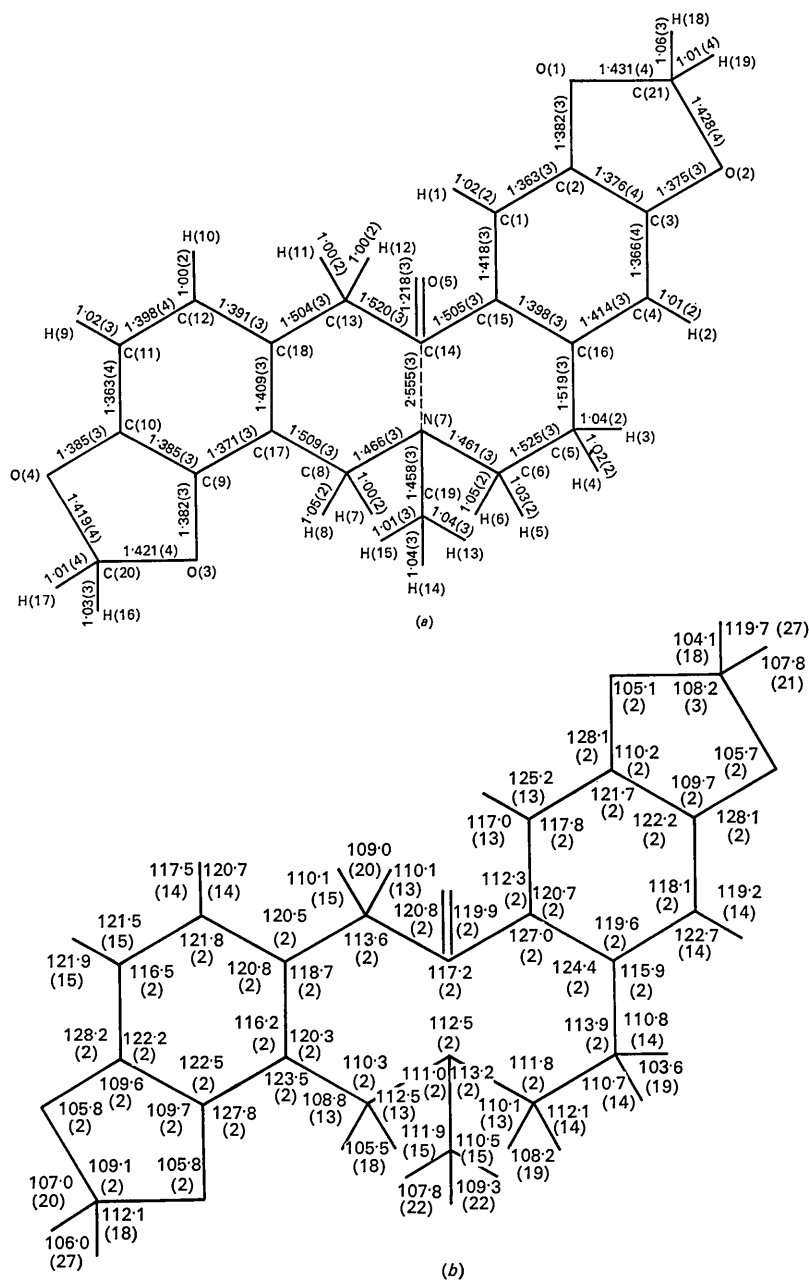


Fig. 4. Bond lengths (Å) and angles (°), and their e.s.d.'s in parentheses. The e.s.d.'s are multiplied by 10² for the C–H bonds, 10³ for the other bonds, and 10 for the angles.

precession photographs available to them. All computations have been carried out with the IBM 360 system using a compatible set of programs written for that computer in FORTRAN IV by Ahmed, Hall, Pippy & Huber (1966). Grateful acknowledgement is made to Mrs M. E. Pippy of this laboratory for assistance with the computations, and to the staff of the NRC Computation Centre for their cooperation.

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The Crystal Structure of Cryptopine, C₂₁H₂₃O₅N

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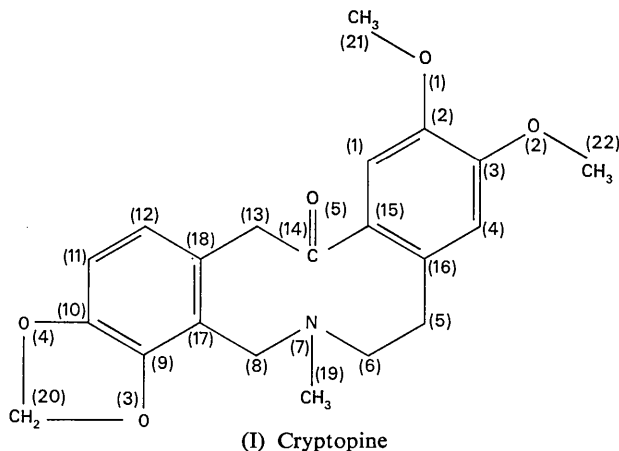
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The structure of cryptopine, C₂₁H₂₃O₅N, has been determined by the symbolic addition procedure, and has been refined by least squares to an *R* index of 0.042 for 1948 observed reflexions measured with a scintillation counter. The space group is *P2₁/n* and the cell dimensions are *a* = 14.438, *b* = 8.702, *c* = 14.369 Å, *β* = 91°22'. The presence of two methoxyl groups in the cryptopine molecule in place of one of the five-membered rings of the protopine molecule has resulted in significant differences in the angles at the two oxygen atoms and at four carbon atoms of the benzene ring in that part of the two molecules. The short distance between the C atom of the carbonyl group and the nitrogen atom across the ten-membered ring is 2.581 Å in cryptopine as compared with 2.555 Å in protopine. In both structures, the nitrogen atom is well buried in the molecule and has the three C-N bonds and the short C...N distance across the ten-membered ring tetrahedrally arranged around it. The carbon-oxygen double bond has ketonic character and is well off the planes of the aromatic rings.

Introduction

Cryptopine, C₂₁H₂₃O₅N, is the dimethyl analogue of protopine, C₂₀H₁₉O₅N, whose X-ray crystal structure analysis has been reported by Hall & Ahmed (1968). Reference should be made to the introduction of that paper where the purpose of the X-ray analysis of these two compounds has been discussed. The schematic formula of cryptopine is shown in (I), and the numbering of the atoms is basically the same as that adopted for protopine for ease of comparison of the two molecules.



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